

Summative Assessment – Semester I**Chemistry**

Class: 8

Max Mks: 45

Time: 2 hrs

Chapters included are Hydrogen, Carbon and Its Compounds and Structure of Atoms

General Instructions:

All questions are compulsory

Questions 1 to 5 are multiple-choice questions and carry 1 mark each. Choose the appropriate option and write the answer in the answer sheet.

Question 6 to 10 carry 1 mark each and must be answered in 1 or 2 sentences.

Questions 11 to 19 carry 2 marks each must be answered in 2 or 3 sentences.

Questions 20 to 23 carry 3 marks each and must be answered in 5 or 6 sentences.

Question 24 carries 5 marks and must be answered in 7 or 8 sentences.

- In free state, hydrogen is present in:
 - Natural gas
 - Sun
 - Petroleum
 - Mineral salts
- When diamond is heated in vacuum for a very long time, it changes into:
 - Graphite
 - Charcoal
 - Lamp black
 - Carbon dioxide
- Nucleons are:
 - Protons only
 - Neutrons only
 - Protons and neutrons
 - Protons and electrons
- Fullerene molecule [C_{60}] is named after:
 - H.W. Kroto
 - R.F. Curl
 - Robert Buckminster Fuller
 - R.E. Smalley

5. The maximum number of electrons in L shell are:
 - (a) 2
 - (b) 18
 - (c) 32
 - (d) 8
6. What is variable valency?
7. Name two crystalline allotropes of carbon.
8. Hydrogen was discovered by _____ in 1766.
9. Hardest naturally occurring substance is _____.
10. What are isotopes?
11. State a few uses of nuclear reactions.
12. State two uses of lamp black.
13. Name four allotropes of carbon.
14. Statements given below are incorrect. Write the correct statements.
 - (i) Hydrogen is used in the manufacture of fertilizers such as potassium nitrate and calcium nitrate.
 - (ii) Liquid sulphur dioxide is mixed with LPG to detect its presence.
15. Write the correct words for the following statements:
 - (i) Atoms of same element having same atomic number but different mass numbers.
_____.
 - (ii) An assembly of protons and neutrons within an atom. _____.
16. What do you understand by terms
 - (i) Atomic number,
 - (ii) Mass number
17. Define the following :
 - (i) Ignition temperature
 - (ii) Combustion
18. How can fire be controlled? State two ways.
19. Name two metals which react with cold water to form hydrogen. Support your answer with fully balanced chemical equations.
20. State four uses of hydrogen.
21. State four properties of wood charcoal.
22. Draw a diagram showing various zones of candle flame.

23. State any three precautions observed in running nuclear reactors.
24. Briefly describe an experiment to prove that water contains two volumes of hydrogen and one volume of oxygen.