

Summative Assessment – Semester I**Chemistry**

Class: 8

Max Mks: 45

Time: 2 hrs

Chapters included are Hydrogen, Carbon and Its Compounds and Structure of Atoms

General Instructions:

All questions are compulsory

Questions 1 to 5 are multiple-choice questions and carry 1 mark each. Choose the appropriate option and write the answer in the answer sheet.

Question 6 to 10 carry 1 mark each and must be answered in 1 or 2 sentences.

Questions 11 to 19 carry 2 marks each must be answered in 2 or 3 sentences.

Questions 20 to 23 carry 3 marks each and must be answered in 5 or 6 sentences.

Question 24 carries 5 marks and must be answered in 7 or 8 sentences.

1. The product which is not obtained from the fractional distillation of crude petroleum is
 - (a) Kerosene oil
 - (b) Lubricating oil
 - (c) Linseed oil
 - (d) Diesel oil
2. The property due to which gases and liquids accumulate on the outer surface of solids is
 - (a) Absorption
 - (b) Adsorption
 - (c) Adhesion
 - (d) Cohesion
3. The maximum number of electrons in L shell are
 - (a) 2
 - (b) 18
 - (c) 32
 - (d) 8
4. Hydrogen removes metals from
 - (a) Metal substances
 - (b) Metal oxides
 - (c) Metal nitrates
 - (d) Metal sulphides

5. The nuclear fuel in stars is
 - (a) Helium
 - (b) Hydrogen
 - (c) Carbon
 - (d) Carbon and oxygen
6. Name the radio isotope used in the cure of thyroid cancer.
7. Write the content of carbon in anthracite variety of coal.
8. Name the form of carbon used in gas masks.
9. Name a metal that does not react with dilute sulphuric acid.
10. A slow moving particle used for splitting nucleus of an atom of uranium is _____.
11. What do you understand by the term valency and variable valency.
12. State two uses of natural gas.
13. Define combustible substances. Give one example.
14. Every _____ parts of water by weight contain _____ part of hydrogen by weight.
15. Define radioactivity. Give four examples of radioactive elements.
16. State two advantages of biogas.
17. Describe your observation when hydrogen burns in air.
18. Write the properties of isotopes.
19. Write the composition of petroleum gas and natural gas.
20. Write the reactions of hydrogen with chlorine gas and metallic oxides.
21. (i) What is LPG? The LPG is an odorless gas, yet the gas supplied for domestic consumption smells. Explain.
 - (ii) State the advantages of LPG as a domestic fuel.
 - (iii) State one important precaution in using LPG gas as domestic fuel.
22. How modern theory of atom contradicts Dalton's atomic theory?
23. (i) What is petroleum?
 - (ii) Name five major fraction obtained and one use of each when petroleum is refined.
24. Prove that H_2 is a reducing agent.